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Di-μ-chlorido-bis[(2,2′-bibenzimidazole)-chloridocadmium(II)]

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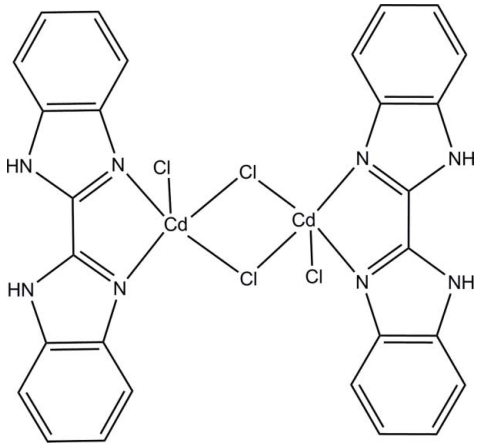
Key indicators: single-crystal X-ray study; T = 293 K; mean σ(C–C) = 0.004 Å; R factor = 0.033; wR factor = 0.060; data-to-parameter ratio = 17.6.

The title binuclear complex, $[Cd_2Cl_4(C_{14}H_{10}N_4)_2]$, was synthesized by the hydrothermal reaction of $CdCl_2$ and the ligand 2,2′-bibenzimidazole. The molecule lies on an inversion center and the metal center displays a strongly distorted trigonal-bipyramidal geometry. The Cd^{II} ions are coordinated by two N atoms from the organic ligand, and by one terminal and two bridging chloride anions. The crystal structure involves intermolecular N—H···Cl hydrogen bonds, resulting in a one-dimensional supramolecular structure.

Related literature

For the synthesis of 2,2′-bibenzimidazole, see: Fieselmann *et al.* (1978). For general properties of Cd^{II}-based complex polymers, see: Meng *et al.* (2004).



Experimental

Crystal data

$[Cd_2Cl_4(C_{14}H_{10}N_4)_2]$ M_r = 835.12
Monoclinic, $C2/c$
a = 11.824 (2) Å
b = 10.784 (2) Å
c = 22.828 (5) Å
β = 91.10 (3)°
V = 2910.1 (10) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 1.86 mm^{-1}
T = 293 (2) K
0.17 × 0.16 × 0.12 mm

Data collection

Rigaku R-AXIS RAPID-S diffractometer
Absorption correction: none
14677 measured reflections
3337 independent reflections
2840 reflections with $I > 2\sigma(I)$
R_{int} = 0.035

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.033
$wR(F^2)$ = 0.060
S = 1.14
3337 reflections
190 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.29 e Å^{-3}
$\Delta\rho_{min}$ = −0.34 e Å^{-3}

Table 1
Selected geometric parameters (Å, °).

Cd1—N4	2.305 (2)	Cd1—Cl1	2.5725 (10)
Cd1—N1	2.338 (2)	Cd1—Cl1^{i}	2.5903 (10)
Cd1—Cl2	2.4602 (8)		
N4—Cd1—Cl2	118.63 (6)	Cl2—Cd1—Cl1	96.65 (3)
N4—Cd1—Cl1	144.04 (6)	N1—Cd1—Cl1^{i}	154.49 (6)

Symmetry code: (i) $-x+2, -y+2, -z+1$.

Table 2
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N3—H16···Cl2^{ii}	0.86	2.39	3.221 (2)	163

Symmetry code: (ii) $x-\frac{1}{2}, y-\frac{1}{2}, z$.

Data collection: *RAPID-AUTO* (Rigaku, 1998); cell refinement: *RAPID-AUTO*; data reduction: *CrystalStructure* (Rigaku/MSC, 2002); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

The author thanks Chifeng University for supporting this work.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BH2211).

References

Fieselmann, B. F., Hendrickson, D. N. & Stucky, G. D. (1978). *Inorg. Chem.* **17**, 2074–2077.
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supplementary materials

Acta Cryst. (2009). E**65**, m37 [**doi:10.1107/S1600536808040968**]

Di-μ-chlorido-bis[(2,2'-bibenzimidazole)chloridocadmium(II)]

G. Liu

Comment

Bibenzimidazole has the potential to function as a bis-bidentate nitrogen ligand by coordinating to metal ions as a chelate. On the other hand, Cd^{II}-containing coordination polymers have attracted much attention as they are able to bond to different donors ligands simultaneously, because of the Cd^{II} large radius. Various coordination modes and potential applications in catalysis, fluorescent materials, NLO materials and so on (Meng *et al.* 2004) have been described. Here we report the crystal structure of the title complex prepared from $CdCl_2$ and bibenzimidazole ligand (see experimental).

As show in Fig. 1, the complex lies on an inversion center, and Cd atoms have strongly distorted trigonal-bipyramidal geometry, being coordinated by two N atoms from the organic ligand, and by one terminal and two bridging Cl^- anions. The two Cd centers are bridged by two chloride ions to give a dinuclear cadmium complex. Intermolecular N—H···Cl hydrogen bonds extend the dinuclear complex to a one dimensional chain in the crystal structure (Fig. 2).

Experimental

A mixture of $CdCl_2$ (0.073 g, 0.40 mmol), bibenzimidazole (0.070 g, 0.30 mmol) and H_2O (10 ml) was placed in a Teflon reactor, then heated to 433 K at 10.8 K/h; after maintaining the reaction at 433 K for three days, it was cooled to 303 K at 10.8 K/h. Crystals suitable for X-ray analysis were obtained.

Refinement

Raw diffraction data were used for refinement, since semi-empirical correction failed to properly correct absorption effects. All H atoms were placed in geometrically idealized positions and constrained to ride on their parent atoms, with N—H = 0.86 Å, C—H = 0.93 Å, and with U_{iso}(H) = 1.2U_{eq}(carrier atom).

Figures

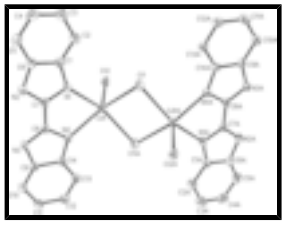

Fig. 1. The structure of the title compound with displacement ellipsoids at the 30% probability level.

Fig. 2. One dimensional chain formed by hydrogen bonds (dashed lines) in the crystal structure of the title compound.

Di-μ-chlorido-bis[(2,2'-bibenzimidazole)chloridocadmium(II)]

Crystal data

$[Cd_2Cl_4(C_{14}H_{10}N_4)_2]$	F_{000} = 1632
M_r = 835.12	D_x = 1.906 Mg m^{-3}
Monoclinic, *C*2/*c*	Mo *K*α radiation λ = 0.71073 Å
Hall symbol: -C 2yc	Cell parameters from 13595 reflections
a = 11.824 (2) Å	θ = 3.1–27.5º
b = 10.784 (2) Å	μ = 1.86 mm^{-1}
c = 22.828 (5) Å	*T* = 293 (2) K
β = 91.10 (3)º	Prism, yellow
V = 2910.1 (10) Å^3	0.17 × 0.16 × 0.12 mm
Z = 4	

Data collection

Rigaku R-AXIS RAPID-S diffractometer	2840 reflections with $I > 2\sigma(I)$
Radiation source: fine-focus sealed tube	R_{int} = 0.035
Monochromator: graphite	θ_{max} = 27.5º
T = 293(2) K	θ_{min} = 3.1º
ω scans	*h* = −15→15
Absorption correction: none	*k* = −14→14
14677 measured reflections	*l* = −29→29
3337 independent reflections	

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)]$ = 0.033	H-atom parameters constrained
$wR(F^2)$ = 0.060	$w = 1/[\sigma^2(F_o^2) + (0.0213P)^2 + 2.9529P]$ where $P = (F_o^2 + 2F_c^2)/3$
S = 1.14	$(\Delta/\sigma)_{max}$ = 0.002
3337 reflections	$\Delta\rho_{max}$ = 0.29 e Å^{-3}
190 parameters	$\Delta\rho_{min}$ = −0.34 e Å^{-3}
Primary atom site location: structure-invariant direct methods	Extinction correction: none

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
C1	0.7642 (2)	0.8420 (3)	0.65287 (12)	0.0352 (6)

C2	0.8455 (3)	0.8628 (3)	0.69666 (14)	0.0488 (8)
H2	0.9096	0.9102	0.6896	0.059*
C3	0.8279 (3)	0.8110 (3)	0.75062 (15)	0.0567 (9)
H3	0.8804	0.8252	0.7807	0.068*
C4	0.7338 (3)	0.7380 (3)	0.76151 (14)	0.0569 (9)
H4	0.7252	0.7037	0.7985	0.068*
C5	0.6533 (3)	0.7156 (3)	0.71877 (13)	0.0523 (9)
H5	0.5909	0.6658	0.7259	0.063*
C6	0.6688 (2)	0.7699 (3)	0.66449 (12)	0.0370 (6)
C7	0.6614 (2)	0.8380 (2)	0.57409 (11)	0.0314 (6)
C8	0.6263 (2)	0.8657 (2)	0.51432 (11)	0.0314 (6)
C9	0.5301 (2)	0.8803 (2)	0.43105 (11)	0.0319 (6)
C10	0.4514 (2)	0.8770 (3)	0.38526 (13)	0.0409 (7)
H10	0.3817	0.8381	0.3889	0.049*
C11	0.4822 (3)	0.9342 (3)	0.33429 (13)	0.0450 (7)
H11	0.4318	0.9343	0.3025	0.054*
C12	0.5875 (3)	0.9926 (3)	0.32871 (13)	0.0441 (7)
H12	0.6050	1.0302	0.2934	0.053*
C13	0.6652 (2)	0.9957 (2)	0.37397 (12)	0.0378 (6)
H13	0.7349	1.0343	0.3700	0.045*
C14	0.6356 (2)	0.9386 (2)	0.42626 (11)	0.0306 (6)
Cd1	0.851704 (16)	1.006362 (18)	0.527622 (9)	0.03487 (7)
Cl1	1.04814 (6)	0.95042 (8)	0.57022 (3)	0.04315 (18)
Cl2	0.84050 (6)	1.21444 (7)	0.57132 (3)	0.04270 (18)
N1	0.75762 (18)	0.8824 (2)	0.59515 (10)	0.0343 (5)
N2	0.60522 (19)	0.7697 (2)	0.61309 (9)	0.0386 (6)
H15	0.5416	0.7327	0.6072	0.046*
N3	0.52742 (18)	0.8349 (2)	0.48775 (9)	0.0347 (5)
H16	0.4728	0.7946	0.5032	0.042*
N4	0.69432 (18)	0.9280 (2)	0.47931 (9)	0.0324 (5)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0327 (15)	0.0355 (15)	0.0373 (15)	0.0013 (11)	0.0013 (12)	0.0007 (12)
C2	0.0377 (17)	0.058 (2)	0.0510 (19)	−0.0033 (15)	−0.0065 (15)	0.0030 (16)
C3	0.051 (2)	0.071 (2)	0.047 (2)	0.0076 (18)	−0.0105 (16)	−0.0022 (18)
C4	0.058 (2)	0.077 (3)	0.0352 (18)	0.0122 (19)	0.0038 (16)	0.0087 (17)
C5	0.0448 (19)	0.067 (2)	0.0455 (19)	−0.0043 (16)	0.0112 (15)	0.0063 (16)
C6	0.0329 (15)	0.0432 (17)	0.0353 (15)	0.0006 (12)	0.0052 (12)	−0.0024 (13)
C7	0.0283 (14)	0.0314 (14)	0.0348 (14)	−0.0073 (11)	0.0049 (11)	−0.0026 (11)
C8	0.0276 (14)	0.0320 (14)	0.0348 (14)	−0.0081 (11)	0.0037 (11)	−0.0045 (11)
C9	0.0307 (14)	0.0312 (14)	0.0338 (14)	−0.0053 (11)	0.0019 (11)	−0.0052 (11)
C10	0.0329 (15)	0.0450 (17)	0.0446 (17)	−0.0079 (13)	−0.0029 (13)	−0.0052 (14)
C11	0.0461 (18)	0.0484 (18)	0.0400 (17)	−0.0064 (14)	−0.0103 (14)	−0.0030 (14)
C12	0.0552 (19)	0.0410 (17)	0.0363 (15)	−0.0103 (14)	0.0022 (14)	0.0006 (13)
C13	0.0383 (15)	0.0364 (15)	0.0389 (15)	−0.0119 (12)	0.0071 (12)	−0.0028 (13)
C14	0.0302 (14)	0.0282 (13)	0.0337 (14)	−0.0057 (11)	0.0044 (11)	−0.0067 (11)

Cd1	0.02621 (11)	0.03603 (12)	0.04251 (12)	−0.01089 (8)	0.00477 (8)	−0.00660 (10)
Cl1	0.0287 (3)	0.0578 (4)	0.0431 (4)	−0.0050 (3)	0.0039 (3)	0.0038 (3)
Cl2	0.0364 (4)	0.0389 (4)	0.0533 (4)	−0.0110 (3)	0.0141 (3)	−0.0121 (3)
N1	0.0256 (11)	0.0377 (13)	0.0396 (13)	−0.0070 (9)	0.0014 (10)	0.0005 (10)
N2	0.0336 (13)	0.0449 (14)	0.0375 (13)	−0.0157 (11)	0.0039 (10)	0.0001 (11)
N3	0.0265 (12)	0.0396 (13)	0.0380 (13)	−0.0141 (10)	0.0039 (10)	−0.0025 (10)
N4	0.0288 (12)	0.0355 (13)	0.0330 (12)	−0.0111 (10)	0.0039 (9)	−0.0015 (10)

Geometric parameters (Å, °)

C1—N1	1.389 (3)	C9—C14	1.403 (3)
C1—C2	1.393 (4)	C10—C11	1.373 (4)
C1—C6	1.399 (4)	C10—H10	0.9300
C2—C3	1.372 (4)	C11—C12	1.403 (4)
C2—H2	0.9300	C11—H11	0.9300
C3—C4	1.389 (5)	C12—C13	1.370 (4)
C3—H3	0.9300	C12—H12	0.9300
C4—C5	1.371 (4)	C13—C14	1.393 (4)
C4—H4	0.9300	C13—H13	0.9300
C5—C6	1.386 (4)	C14—N4	1.389 (3)
C5—H5	0.9300	Cd1—N4	2.305 (2)
C6—N2	1.381 (3)	Cd1—N1	2.338 (2)
C7—N1	1.317 (3)	Cd1—Cl2	2.4602 (8)
C7—N2	1.341 (3)	Cd1—Cl1	2.5725 (10)
C7—C8	1.450 (4)	Cd1—Cl1^{i}	2.5903 (10)
C8—N4	1.327 (3)	Cl1—Cd1^{i}	2.5903 (10)
C8—N3	1.348 (3)	N2—H15	0.8600
C9—N3	1.385 (3)	N3—H16	0.8600
C9—C10	1.387 (4)		
N1—C1—C2	131.0 (3)	C13—C12—C11	121.7 (3)
N1—C1—C6	108.9 (2)	C13—C12—H12	119.1
C2—C1—C6	120.1 (3)	C11—C12—H12	119.1
C3—C2—C1	117.6 (3)	C12—C13—C14	117.4 (3)
C3—C2—H2	121.2	C12—C13—H13	121.3
C1—C2—H2	121.2	C14—C13—H13	121.3
C2—C3—C4	121.8 (3)	N4—C14—C13	130.9 (2)
C2—C3—H3	119.1	N4—C14—C9	109.0 (2)
C4—C3—H3	119.1	C13—C14—C9	120.2 (3)
C5—C4—C3	121.3 (3)	N4—Cd1—N1	73.49 (8)
C5—C4—H4	119.3	N4—Cd1—Cl2	118.63 (6)
C3—C4—H4	119.3	N1—Cd1—Cl2	102.95 (6)
C4—C5—C6	117.4 (3)	N4—Cd1—Cl1	144.04 (6)
C4—C5—H5	121.3	N1—Cd1—Cl1	93.11 (6)
C6—C5—H5	121.3	Cl2—Cd1—Cl1	96.65 (3)
N2—C6—C5	132.8 (3)	N4—Cd1—Cl1^{i}	91.81 (6)
N2—C6—C1	105.5 (2)	N1—Cd1—Cl1^{i}	154.49 (6)
C5—C6—C1	121.7 (3)	Cl2—Cd1—Cl1^{i}	102.39 (3)
N1—C7—N2	113.1 (2)	Cl1—Cd1—Cl1^{i}	86.78 (3)

N1—C7—C8	119.9 (2)	Cd1—Cl1—Cd1^{i}	93.22 (3)
N2—C7—C8	127.0 (2)	C7—N1—C1	105.3 (2)
N4—C8—N3	112.6 (2)	C7—N1—Cd1	112.71 (17)
N4—C8—C7	120.4 (2)	C1—N1—Cd1	141.87 (18)
N3—C8—C7	127.1 (2)	C7—N2—C6	107.1 (2)
N3—C9—C10	131.9 (2)	C7—N2—H15	126.4
N3—C9—C14	105.5 (2)	C6—N2—H15	126.4
C10—C9—C14	122.5 (3)	C8—N3—C9	107.4 (2)
C11—C10—C9	116.2 (3)	C8—N3—H16	126.3
C11—C10—H10	121.9	C9—N3—H16	126.3
C9—C10—H10	121.9	C8—N4—C14	105.6 (2)
C10—C11—C12	122.0 (3)	C8—N4—Cd1	112.97 (16)
C10—C11—H11	119.0	C14—N4—Cd1	140.52 (16)
C12—C11—H11	119.0		
N1—C1—C2—C3	179.2 (3)	C6—C1—N1—C7	1.1 (3)
C6—C1—C2—C3	−0.3 (5)	C2—C1—N1—Cd1	−2.8 (5)
C1—C2—C3—C4	1.3 (5)	C6—C1—N1—Cd1	176.7 (2)
C2—C3—C4—C5	−0.8 (5)	N4—Cd1—N1—C7	−4.14 (18)
C3—C4—C5—C6	−0.8 (5)	Cl2—Cd1—N1—C7	112.26 (18)
C4—C5—C6—N2	−178.4 (3)	Cl1—Cd1—N1—C7	−150.18 (18)
C4—C5—C6—C1	1.9 (5)	Cl1^{i}—Cd1—N1—C7	−61.1 (2)
N1—C1—C6—N2	−0.7 (3)	N4—Cd1—N1—C1	−179.5 (3)
C2—C1—C6—N2	178.8 (3)	Cl2—Cd1—N1—C1	−63.1 (3)
N1—C1—C6—C5	179.1 (3)	Cl1—Cd1—N1—C1	34.5 (3)
C2—C1—C6—C5	−1.3 (4)	Cl1^{i}—Cd1—N1—C1	123.5 (3)
N1—C7—C8—N4	5.0 (4)	N1—C7—N2—C6	0.7 (3)
N2—C7—C8—N4	−175.5 (3)	C8—C7—N2—C6	−178.9 (3)
N1—C7—C8—N3	−175.0 (3)	C5—C6—N2—C7	−179.7 (3)
N2—C7—C8—N3	4.6 (5)	C1—C6—N2—C7	0.0 (3)
N3—C9—C10—C11	178.9 (3)	N4—C8—N3—C9	−0.1 (3)
C14—C9—C10—C11	0.0 (4)	C7—C8—N3—C9	179.8 (3)
C9—C10—C11—C12	0.1 (5)	C10—C9—N3—C8	−178.9 (3)
C10—C11—C12—C13	−0.1 (5)	C14—C9—N3—C8	0.1 (3)
C11—C12—C13—C14	−0.2 (4)	N3—C8—N4—C14	0.1 (3)
C12—C13—C14—N4	−178.9 (3)	C7—C8—N4—C14	−179.9 (2)
C12—C13—C14—C9	0.4 (4)	N3—C8—N4—Cd1	171.44 (17)
N3—C9—C14—N4	0.0 (3)	C7—C8—N4—Cd1	−8.5 (3)
C10—C9—C14—N4	179.1 (2)	C13—C14—N4—C8	179.3 (3)
N3—C9—C14—C13	−179.5 (2)	C9—C14—N4—C8	0.0 (3)
C10—C9—C14—C13	−0.3 (4)	C13—C14—N4—Cd1	11.9 (5)
N4—Cd1—Cl1—Cd1^{i}	88.65 (9)	C9—C14—N4—Cd1	−167.4 (2)
N1—Cd1—Cl1—Cd1^{i}	154.46 (6)	N1—Cd1—N4—C8	6.56 (18)
Cl2—Cd1—Cl1—Cd1^{i}	−102.11 (3)	Cl2—Cd1—N4—C8	−89.44 (19)
Cl1^{i}—Cd1—Cl1—Cd1^{i}	0.0	Cl1—Cd1—N4—C8	78.4 (2)
N2—C7—N1—C1	−1.1 (3)	Cl1^{i}—Cd1—N4—C8	165.38 (18)
C8—C7—N1—C1	178.5 (2)	N1—Cd1—N4—C14	173.4 (3)
N2—C7—N1—Cd1	−178.16 (18)	Cl2—Cd1—N4—C14	77.4 (3)

C8—C7—N1—Cd1	1.4 (3)	Cl1—Cd1—N4—C14	−114.8 (3)
C2—C1—N1—C7	−178.4 (3)	Cl1^{i}—Cd1—N4—C14	−27.8 (3)

Symmetry codes: (i) −*x*+2, −*y*+2, −*z*+1.

Hydrogen-bond geometry (Å, °)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N3—H16···Cl2^{ii}	0.86	2.39	3.221 (2)	163

Symmetry codes: (ii) *x*−1/2, *y*−1/2, *z*.

Fig. 1

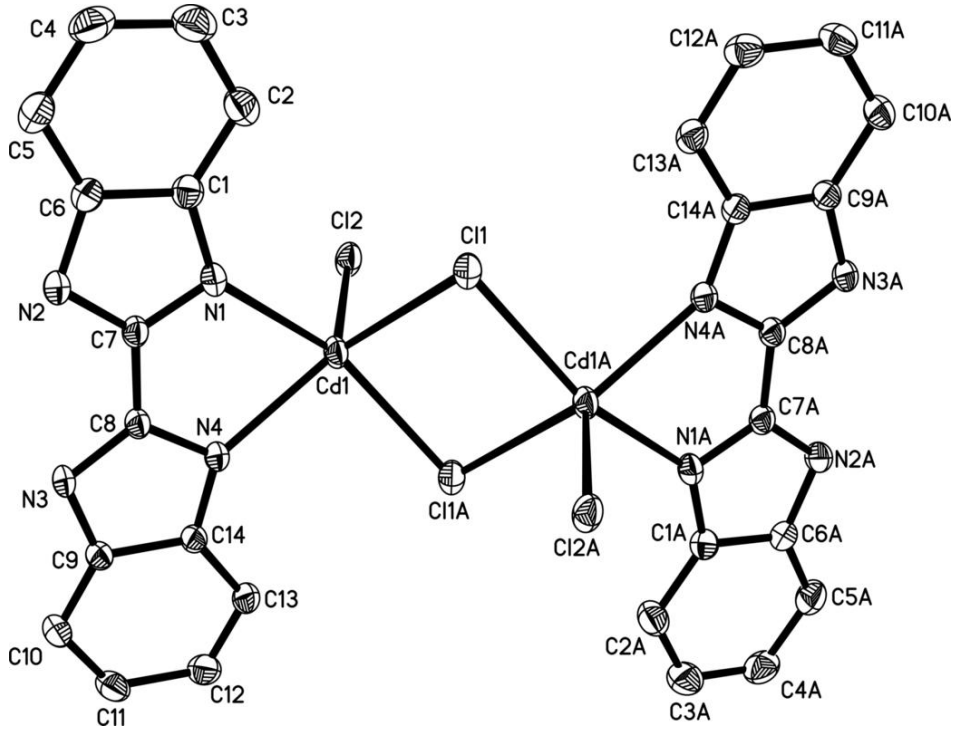
C4
C3
C2
C5
C1
C6
N2
C7
N1
Cl2
Cl1
Cd1
C8
N4
N3
C14
C9
C13
C10
C12
C11
Cl1A
Cl2A
Cd1A
C12A
C11A
C10A
C13A
C14A
C9A
N3A
N4A
C8A
C7A
N1A
N2A
C1A
C6A
C5A
C2A
C4A
C3A

Fig. 2